

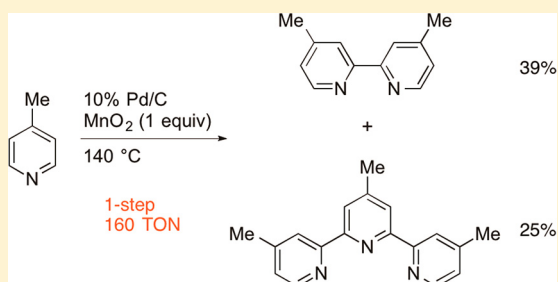
4,4',4''-Trimethyl-2,2':6',2''-terpyridine by Oxidative Coupling of 4-Picoline

Michael T. Robo,[†] Michael R. Prinsell,[‡] and Daniel J. Weix*

Department of Chemistry, University of Rochester, Rochester, New York 14627-0216, United States

S Supporting Information

ABSTRACT: Alkylated terpyridine ligands are an increasingly important component of catalysis and dyes but are costly because their synthesis is challenging and often low-yielding. We report an improved method for the Pd/C-catalyzed dehydrogenative coupling of 4-picoline to form the bi- and terpyridine. The addition of MnO₂ improves the yield of the reaction, making the reaction useful on a large scale (up to 200 mmol). The use of Pd(OAc)₂ or Pd/C/pivalic acid leads to the selective formation of bipyridine.



Although terpyridines are increasingly used in catalysis and in light-harvesting materials, their synthesis remains challenging.¹ For example, 4,4',4''-trimethylterpyridine (**1a**) has found use as a ligand in dye-sensitized nanocrystalline solar cells² and as a precursor for bisterpyridylalkanes (Figure 1).³ It

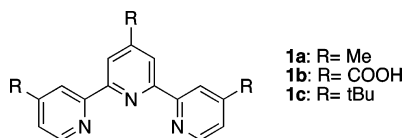


Figure 1. Terpyridine ligands.

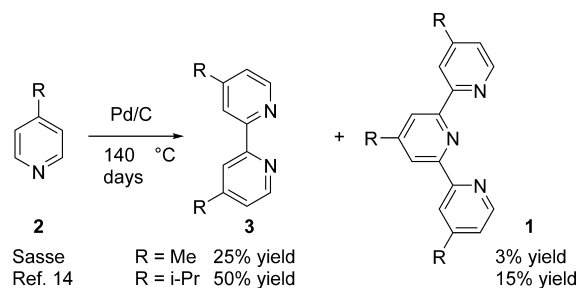
is also a precursor to the tricarboxylate (**1b**),⁴ which has also proven to be useful in dye-sensitized solar cells.⁵ In addition, commercially available tri-*tert*-butylterpyridine **1c** was shown by Vicic to be a useful ligand for forming Csp³–Csp³ bonds in Negishi cross-coupling,⁶ and we found it to be useful in reductive cross-electrophile coupling.⁷ On the other hand, most terpyridines are costly and only available in small quantities (1–5 g).⁸ For example, ligand **1c** is only available from one large chemical supplier,⁹ and at times it has been back ordered for months. We had previously used the procedure of Ben Hadda and Le Bozec to synthesize **1c**¹⁰ but found the yields to be variable. Herein, we demonstrate a reliable, scalable procedure for the synthesis of **1a** and demonstrate that it is a reasonable replacement for **1c** in cross-electrophile coupling.

Most methods for the synthesis of terpyridines require functionalization of the starting pyridine, adding several steps for each ligand. Condensation to form one or more of the pyridyl groups,^{1a,11} cross-coupling of 2-pyridylmetal reagents with 2,6-dihalopyridines,¹² and C–H arylation of pyridine and bipyridine *N*-oxides with 2-halopyridines have all been reported.¹³

The most direct route to 4,4',4''-trisubstituted terpyridines is the dehydrogenative trimerization of alkylpyridines (Scheme

1), first reported by Rosevear and Sasse. They found that a number of alkylpyridines, when refluxed in the presence of

Scheme 1. Previous Reports on Dehydrogenative Coupling of Alkylpyridines



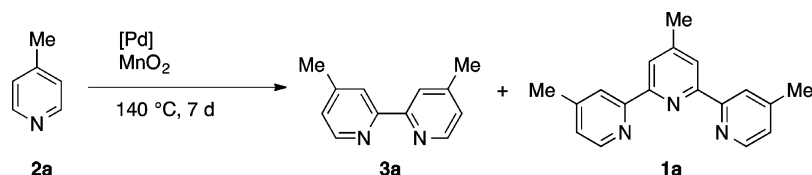
palladium on carbon (Pd/C), gave mixtures of the corresponding bipyridines (10–50% yield) and terpyridines (0.5–15% yield).¹⁴ The highest yield was reported with 4-isopropylpyridine. Based on this work, Ben Hadda and Le Bozec developed their synthesis of **1c** from 4-*tert*-butylpyridine.¹⁰ Further effort on developing this reaction has focused on improving the yield of bipyridine and the range of functional groups tolerated.¹⁵ Recent studies by Åkermark and Hagelin-Weaver demonstrated that a terminal oxidant is needed¹⁶ and that air can serve as the oxidant.¹⁷ Bipyridines, but not terpyridines, can also be made using nickel catalysis.¹⁸

Our goal was to develop a one-step synthesis of a trialkylterpyridine ligand that would scale and had the potential to be low cost. We chose to focus on terpyridine **1a** due to its expected electronic similarity to **1c**, the lower cost of 4-picoline compared to 4-*tert*-butylpyridine,¹⁹ and the established utility of

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Table 1. Optimization of Reaction Conditions



entry ^a	[Pd] source	[Pd] (mol %)	MnO ₂ (equiv)	remaining 2a (%)	yield of 3a (%)	yield of 1a (%)	TON ^b	3a:1a ^c
1	10% Pd/C	0.18	0.25	81	18	1	111	26.9
2	10% Pd/C	0.18	0.5	73	26	1	156	35.1
3	10% Pd/C	0.18	1.0	61	38	3	244	19.1
4	10% Pd/C	0.18	2.0	64	28	1	167	28.3
5	10% Pd/C	0.37	0.5	48	43	6	149	11.7
6	10% Pd/C	0.73	0.5	31	27	10	64	4.0
7 ^d	10% Pd/C	0.73	1.0	6	56	29	156	2.9
8	10% Pd/C	0.73	0	78	11	0	15	NA
9	10% Pd/C	0	1.0	95	0	0	NA	NA
10 ^{d,e}	10% Pd/C	0.73	1.0	2	53	30	155	2.6
11 ^{e,f}	10% Pd/C	0.75	1.0	9	46 (39) ^f	37 (25) ^f	160	1.9
12 ^d	30% Pd/C	2.19	1.0	74	14	3	9	7.3
13 ^e	Pd(OAc) ₂	3.0	1.0	30	50	1	17	98.4
14 ^{e,g}	10% Pd/C	0.73	1.0	38	51	0	70	NA

^aReaction set up in glovebox in 1-dram vial with 20.55 mmol of 4-picoline, Pd/C, and MnO₂. This vial was sealed and stirred at 140 °C for 7 d. Yields are corrected GC yields vs dodecane as an internal standard. Isolated yields in parentheses. ^bTurnover number calculated as number of C–C bonds formed/Pd atom. ^cThe ratio of 3a to 1a, not the ratio of yields of 3a to 1a. For example, 50% yield of 3a and a 50% yield of 1a would yield 5.14 mmol of 3a and 3.43 mmol of 1a; thus, the ratio of 3a:1a would be 1.5:1. ^dRan for 6 days. ^eSet up using an oven-dried vial on the benchtop. ^fUsed 1.95 mL (20 mmol) of 4-picoline. ^gWith 20 mol % of 2,2-dimethylpropionic acid.

1a in light harvesting. Unfortunately, the highest reported yield for the dehydrogenative synthesis of 1a from 4-picoline was 3% (Scheme 1).¹⁴

In preliminary studies, the major challenges were low conversion of picoline and a high ratio of bipyridine (3a) to our desired terpyridine (1a). In order to improve the turnover numbers of the catalyst, we sought a chemical oxidant to regenerate palladium(II) oxide, since oxidants are typically required for dehydrogenative coupling reactions.²⁰ Although Hagelin-Weaver found that air was suitable for their catalysts,¹⁷ we did not see any improvement in yields with commercial Pd/C. A variety of solid oxidants were tested (CuCl₂,²¹ FeCl₃, benzoquinone,²² DDQ, and Oxone²³), but only unactivated MnO₂ was found to produce significant yields of terpyridine 1a and bipyridine 3a (Table 1), which we attribute to the reoxidation of Pd(0) to Pd(II). While activated MnO₂ is known to oxidize Pd(0) to Pd(II),²⁴ use of commercial activated MnO₂ in this procedure yielded only starting material.²⁵ Increasing the oxidant loading to 1 equiv increased the turnover number (TON) and slightly improved the ratio of 3a:1a (entries 1–3). Beyond 1 equiv, yields dropped as stirring the solution became difficult (entry 4). Further improvements in the conversion of picoline to coupled products and the ratio of 3a:1a were obtained by increasing the catalyst loading (entry 2 vs entries 5 and 6).

These two effects were cumulative, and reaction of picoline with 1 equiv of MnO₂ at a loading of 0.73 mol % of Pd resulted in nearly complete conversion and a 2.9:1 ratio of 3a:1a (entry 7). The same reaction conducted without MnO₂ provided no terpyridine and only a small amount of bipyridine (entry 8), confirming the positive role of the oxidant. Similarly, Pd is required for coupling to occur (entry 9). The reaction could be set up on the benchtop, as long as all materials were rigorously dry (entries 10 and 11), and the products could be isolated by

sublimation in reasonable yield. The remaining material contained higher oligopyridines up to the sexipyridine (determined by LC–MS analysis).

Most of the other Pd sources we tested²⁶ were found to be unsuitable catalysts with two exceptions: 30% Pd/C (entry 12) and Pd(OAc)₂ (entry 13). Although Pd(OAc)₂ is a well-known catalyst for the oxidative coupling of arenes and C–H functionalization,²⁰ in our reactions Pd metal precipitates within the first few hours. Even after this plating out, however, product continues to be formed. For both the Pd/C- and the Pd(OAc)₂-catalyzed reactions, it is unclear whether heterogeneous catalysis, as suggested by Hagelin-Weaver's work,^{16,17} or homogeneous catalysis²⁰ is at work. The Pd(OAc)₂-catalyzed reaction was highly selective for bipyridine formation (3a), which was different from the Pd/C-catalyzed reactions (entry 13 vs entries 1–12). However, a similar selectivity was observed when 2,2-dimethylpropionic acid was added to a reaction catalyzed by 10% Pd/C (entry 14), suggesting that the carboxylate ligands alter the selectivity and that the two catalysts operate in a similar fashion.

A study of the amount of product formed over time (Figure 2) shows that bipyridine is produced continuously until day 4 or 5 and terpyridine production is minimal until about day 4. We found that the addition of bipyridine 3a does not significantly increase the amount of terpyridine 1a produced, consistent with the observation of Sasse.¹⁴ At this time, it is unclear why a mixture of 2a and 3a does not produce more 1a.

To test the scalability of this heterogeneous reaction, we conducted a synthesis of 1a from 19.5 mL (200 mmol) of 4-picoline (2a) on the benchtop using a mechanical stirrer, a three-neck flask, and a reflux condenser (Scheme 2). The reaction scaled well, affording gram quantities of 1a and 3a in a single step, albeit with a long reaction time.

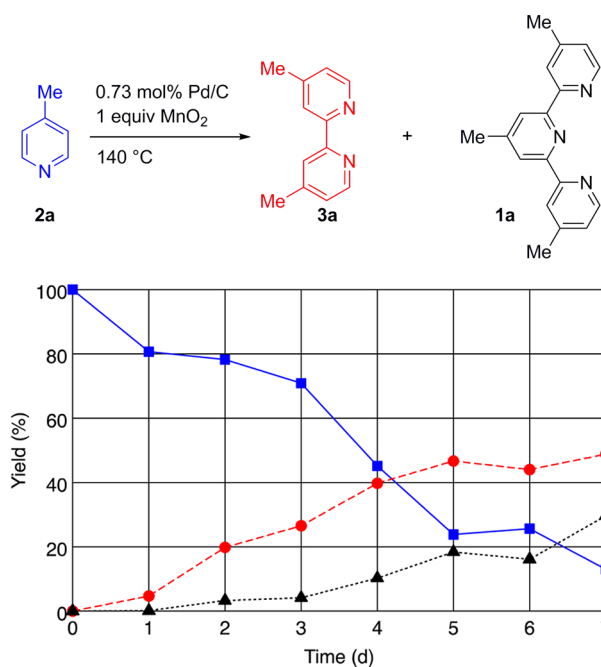
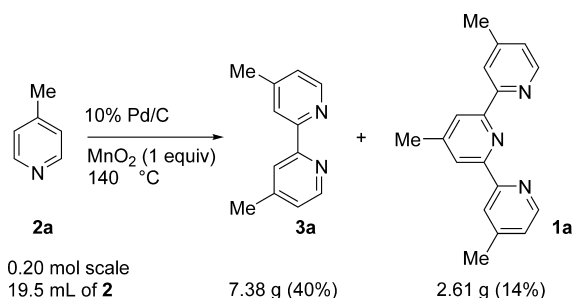


Figure 2. Bipyridine and terpyridine formation over time. Legend: **2a** (blue boxes, solid line), **3a** (red circles, dashed line), and **1a** (black triangles, dotted line). Procedure: 12 reactions were set up according to Table 2, entry 8. Each day, two reactions were halted, except days 6 and 7, where one reaction was halted.

Scheme 2. Large-Scale Synthesis of Trimethylterpyridine **1a**



We also briefly examined the substrate scope of this procedure. Overall, the scope is promising and improved compared to literature reports, but conversions, selectivities, and yields were not as high as with 4-picoline (Table 2). As expected, 3-picoline only produced trace amounts of terpyridine **2d**, but a 31% isolated yield of bipyridine **3d** was obtained. This compares favorably with the literature yields of 36% (Raney Ni)²⁷ and 0.5% (Pd/C).²⁸ The reaction of the electron-rich pyridine, 4-(dimethylamino)pyridine (**2f**), afforded a 40% yield of bipyridine **3f** along with a 16% yield of terpyridine **1f**. While the synthesis of bipyridine **3f** by ruthenium catalysis²⁹ is known in 32% yield, trimerization to form **1f** has not been previously reported. The slightly electron withdrawing 4-phenylpyridine (**2e**) (Table 2, entries 2 and 3) could be dimerized to **3e** in 49% yield along with an 11% yield of the known terpyridine **1e**,³⁰ but more electron-poor pyridines were largely unreactive. The dimerization of 4-phenylpyridine to bipyridine **3e** over Raney Ni was previously reported in 5% yield,²⁷ but the trimerization has not been reported. Finally, we found 4-*tert*-butylpyridine **2c** to be relatively unreactive for reasons we are unable to explain at this time (entry 4). As an alternative, a reasonable yield of 4,4'-

Table 2. Substrate Scope

entry ^a	substrate	2 (%)	yield of 3 (%)	yield of 1 (%)	3 : 1
1 ^{b,c}	R = 3-Me, 2d	34	53 (31)	1	53:1
2 ^{d,e}	R = 4-Ph, 2e	22	49	11	4.5:1
3 ^d	R = 4-NMe ₂ , 2f	40	40	16	2.5:1
4 ^f	R = 4- <i>t</i> -Bu, 2c	91	8	0.7	11:1
5	R = 4-SiMe ₃ , 2g	30	42 (42)	2	21:1

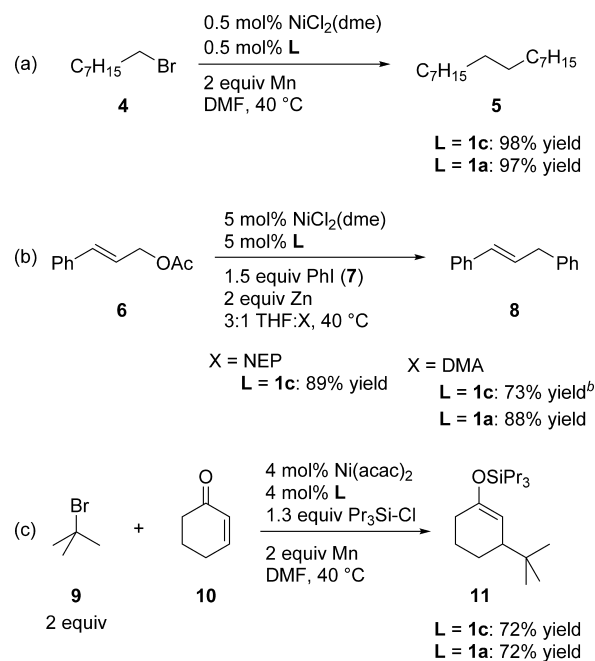
^aSet up on the benchtop in an oven-dried 1-dram vial with substrate, 0.75 mol % of 10% Pd/C, and 100 mol % of MnO₂ at 190 °C. Yields are assay yields based upon GC areas % data. Isolated yields are given in parentheses. ^bCorrected GC yield. ^cRun at 140 °C. ^dYield determined by ¹H NMR. ^eAverage of 2 runs. ^f1.5 mol % Pd/C was used.

bis(trimethylsilyl)-2,2'-bipyridine (**3g**) could be obtained (entry 5). The only previous synthesis of this bipyridine was from the dibromobipyridine in 12% yield.³¹ Due to the moderate selectivity and conversions obtained in entries 2–4, we did not isolate the products of these reactions.

Finally, we tested **1a** as a replacement for **1c** in our published cross-electrophile coupling reactions (Scheme 3). In each case, reactions with **1a** provided results equal to or better than our published results.⁷

In summary, an economical, one-step method to make 4,4',4''-trimethyl-2,2':6',2''-terpyridine (**1a**) has been developed. The use of MnO₂ as an oxidant was found to dramatically

Scheme 3. Performance of **1a** in Reductive Cross-Electrophile Coupling Reactions^{a,b}



^aYields are isolated yields of purified product. ^bFrom ref 7c, GC yield, corrected.

increase the conversion of the oxidative coupling and enhance the amount of terpyridine produced, while an increased catalyst loading further improved the ratio of terpyridine to bipyridine. Unactivated MnO₂ was the only oxidant studied that improved TON and it might prove generally useful in other palladium-catalyzed oxidative chemistry.^{20,24} Finally, ligand **1a** is a suitable replacement for **1c** in cross-electrophile and dimerization chemistry, suggesting that **1a** will be generally useful in catalysis.

EXPERIMENTAL SECTION

General. MnO₂ (>99%, not “activated” grade) and palladium on carbon (Pd/C) (10 wt %, matrix carbon, dry support, Sigma-Aldrich) were used as received. 4-Picoline was distilled from KOH prior to use. 4-(Trimethylsilyl)pyridine (**2g**) was prepared according to the literature procedure.³²

Methods. NMR chemical shifts are reported in ppm, referenced to the residual solvent peak of CDCl₃ ($\delta = 7.26$ ppm for ¹H or $\delta = 77.16$ ppm for ¹³C) as an internal standard, and *J* values are given in hertz. NMR spectra were recorded at 400.13 or 500.13 MHz proton NMR frequency. GC analyses were performed on a DB-5 column (20 m × 180 μ m × 0.18 μ m) with an FID detector, and with H₂ as carrier gas. General method: 1 μ L injection of sample, injection temp of 300 °C, 100:1 split ratio, and initial inlet pressure was 20.3 psi but varied as the column flow was held constant at 1.8 mL/min for the duration of the run. Initial oven temperature of 50 °C was held for 0.46 min followed by a temperature ramp up to 300 °C at 65 °C/min, and finally the temperature was held at 300 °C for 2.69 min. Total run time was ~7 min. FID temperature was 325 °C. Chromatography was performed on SiO₂ (silica gel 60, particle size 0.040–0.063 mm) using standard flash techniques. Products were visualized by either UV light or phosphomolybdic acid.

General Procedures. General Procedure A: Glove Box. Inside a nitrogen glovebox, a 1-dram vial was charged with 10% Pd/C (160 mg, 0.73 mol %), MnO₂ (1.79 g, 20.6 mmol), 4-picoline (2.00 mL, 1.91 g, 20.6 mmol), and a polytetrafluoroethylene-coated (PTFE) magnetic stirbar. The vial was then capped with a screw cap fitted with a PTFE-faced silicone septum, removed from the glovebox and then shaken for 15 s before being stirred (1200 rpm) at 140 °C for 7 d. After 7 d, the reaction mixture was carefully opened and poured hot into a 20 mL scintillation vial. The 1 dram vial was washed multiple times with CH₂Cl₂ (~15 mL in total), and these washings were added to the scintillation vial, along with any remaining solids. The resulting mixture was sonicated for 15 min to ensure all soluble materials were fully dissolved before it was filtered through a plug of diatomaceous earth (1 in. wide, ~1 in. high) in an 18 mL disposable polyethylene fritted filter funnel. The diatomaceous earth pad was washed with ~10 mL CH₂Cl₂. Samples from this mixture were used for GC analysis. **General Procedure B: Benchtop.** As procedure A, except the reaction as set up on the benchtop with no precautions to exclude air. However, to exclude moisture, oven-dried glassware was used. **General Procedure C: Isolated Yield.** As procedure B, except ~100 mL of CH₂Cl₂ was used in filtration, and 159.6 mg of 10% Pd/C (0.75 mol %), 1.74 g of MnO₂ (20 mmol), and 1.95 mL of 4-picoline (1.86 g, 20 mmol) were used.

4,4'-Dimethyl-2,2'-bipyridine (3a) (CAS No. 1134-35-6).³³ General procedure C was followed. The crude solid was purified by sublimation (150 °C/20 mTorr, 36 h). The sublimate was removed from the coldfinger and sublimed again (100 °C/20 mTorr, 8 h). The sublimate of this second sublimation was a white solid whose spectra matched the literature for **3a** (0.71 g, 3.85 mmol, 39% yield). ¹H NMR (500 MHz; CDCl₃): δ 8.54 (d, *J* = 4.9 Hz, 2H), 8.23 (s, 2H), 7.13 (d, *J* = 4.6 Hz, 2H), 2.44 (s, 6H). ¹³C NMR (126 MHz; CDCl₃): δ 156.0, 148.9, 148.1, 124.6, 122.0, 21.2. Compound **3a** was purified further by recrystallization in ethyl acetate (recrystallized yield 0.40 g, 2.16 mmol, 22% yield). Mp: 166–171 °C (lit.³² mp 171.9–172.4 °C). Anal. Calcd for C₁₂H₁₂N₂: C, 78.23; H, 6.57; N, 15.21. Found: C, 77.88; H, 6.52; N, 14.96

4,4',4''-Trimethyl-2,2':6',2''-terpyridine (1a) (CAS No. 33354-75-5).^{12a} The white material remaining after the second sublimation of **3a** was pure **1a** (0.47 g, 1.69 mmol, 25% yield). The ¹H NMR spectrum was in good agreement with the literature report. The ¹³C NMR spectrum was consistent with the structure and the symmetry of **1a** but did not match the spectrum previously reported (ref 12a contains an extra peak at 123.5 ppm and is missing the peak at 149.1 ppm). Mp: 180–182 °C (lit.^{12a} 186 °C). ¹H NMR (500 MHz; CDCl₃): δ 8.59 (d, *J* = 5.0 Hz, 2H), 8.44 (s, 2H), 8.30 (s, 2H), 7.18 (d, *J* = 5.1 Hz, 2H), 2.55 (s, 3H), 2.53 (s, 6H). ¹³C NMR (126 MHz; CDCl₃): δ 156.2, 155.5, 149.1, 148.9, 148.0, 124.7, 122.08, 122.02, 21.38, 21.35. Anal. Calcd for C₁₈H₁₇N₃: C, 78.52; H, 6.22; N, 15.26. Found: C, 78.18; H, 6.24; N, 15.12.

Large-Scale Procedure. 200 mmol Scale. While hot, a reflux condenser, mechanical stirring rod, and plug were attached to a three-neck round-bottom flask. A N₂ inlet and outlet were then attached to the reflux condenser, and the apparatus was cooled under N₂. The stirring rod was then attached to an air-powered mechanical stirrer. Pd/C (10%, 1.60 g, 0.75 mol %), MnO₂ (17.4 g, 200 mmol), and 4-picoline (**2a**) (19.5 mL, 200 mmol, distilled over KOH) were added to the apparatus under positive N₂ pressure while stirring. This mixture was then heated to reflux. After 7 days, 25 mL of mesitylene was added to the mixture, and then the mixture was allowed to cool. This mixture was filtered through a 150 mL coarse frit with 1 in. of diatomaceous earth on top of 2 cm of sand with 1 L of CH₂Cl₂. The CH₂Cl₂ was removed by rotary evaporation, and the remaining **2a** and mesitylene was removed on a high-vacuum line (~50 mTorr).

4,4'-Dimethyl-2,2'-bipyridine (3a) (CAS No. 1134-35-6).³² The crude solid was sublimed for 15 h (100 °C/50 mTorr). The sublimate was isolated as a white solid whose spectra matched the literature spectra for **3a** (7.38 g, 40.1 mmol, 40% yield).

4,4',4''-Trimethyl-2,2':6',2''-terpyridine (1a) (CAS No. 33354-75-5).^{12a} The solid remaining after the sublimation above was then sublimed for 21 h (135 °C/50 mTorr). The sublimate was isolated as a white solid whose spectra matched the literature spectra for **1a** (2.61 g, 9.48 mmol, 14% yield).

5,5'-Dimethyl-2,2'-bipyridine (3b) (CAS No. 1762-34-1).³² General procedure C was employed with 3-picoline (**2b**) (1.95 mL, 20 mmol). Two 20 mmol runs were combined for purification and isolation. After 7 days, the remaining starting material was removed by rotary evaporation, and then the crude solid was purified by sublimation (80 °C/20 mTorr, 12 h). The sublimate was isolated as a light brown solid whose spectra matched the literature spectra for **3b** (1.12 g, 31% yield overall). Mp: 108–112 °C (lit.³² mp 113.7–114.1 °C). ¹H NMR (400 MHz; CDCl₃): δ 8.48 (s, 2H), 8.23 (d, *J* = 8.1 Hz, 2H), 7.59 (dd, *J* = 8.0, 1.6 Hz, 2H), 2.37 (s, 6H). ¹³C NMR (126 MHz; CDCl₃): δ 153.8, 149.5, 137.4, 133.0, 120.3, 18.4. Anal. Calcd for C₁₂H₁₂N₂: C, 78.23; H, 6.57; N, 15.21. Found: C, 77.76; H, 6.53; N, 15.03.

4,4'-Bis(trimethylsilyl)-2,2'-bipyridine (3g) (CAS No. 372805-52-1). General procedure C was followed with 4-(trimethylsilyl)pyridine (**2g**) (0.48 g, 3.2 mmol) using 280 mg MnO₂ (3.2 mmol) and 25.6 mg 10% Pd/C (0.75 mol %). After 7 days, the remaining starting material was evaporated and the crude solid was purified by sublimation (120 °C/70 mTorr, 1 h) to afford the product as an off-white solid (0.20 g, 42% yield). ¹H NMR (500 MHz; CDCl₃): δ 8.65 (d, *J* = 4.7 Hz, 2H), 8.48 (s, 2H), 7.41 (d, *J* = 4.7 Hz, 2H), 0.34 (s, 18H). ¹³C NMR (126 MHz; CDCl₃): δ 155.2, 151.3, 148.0, 128.1, 125.7, -1.6.

Hexadecane (5) (CAS No. 544-76-3). Product **5** was prepared according to the literature procedure on a 2 mmol scale except that ligand **1a** was utilized in place of **1c**. Hexadecane (226 mg, 100% yield) was isolated as a clear liquid whose spectra matched those reported. The reaction was repeated, and 215 mg (95% yield) of **5** was obtained.

(E)-Prop-1-ene-1,3-diylidibenzene (8) (CAS No. 3412-44-0).^{7c} Product **8** was prepared according to the literature procedure except that ligand **1a** was utilized in place of **1c**. After purification by column chromatography (100% pentane), 171 mg (88% yield) of **8** was isolated as a clear liquid. Spectra matched those previously reported in the literature.

[[3-*tert*-Butylcyclohex-1-en-1-yl)oxy]tripropylsilane (**11**) (CAS No. 1286212-20-1).⁷⁶ Product **11** was prepared according to the literature procedure on a 1.00 mmol scale except that ligand **1a** was utilized in place of **1c**. The compound was isolated as a clear liquid (223 mg, 72% yield) whose spectra matched the literature report.

■ ASSOCIATED CONTENT

■ Supporting Information

Copies of ¹H and ¹³C NMR spectra. This material is available free of charge via the Internet at <http://pubs.acs.org>.

■ AUTHOR INFORMATION

Corresponding Author

*E-mail: daniel.weix@rochester.edu.

Present Addresses

[†]Department of Chemistry, University of Michigan, Ann Arbor, MI 48109-1055.

[‡]Department of Chemistry, Trinity University, San Antonio, TX 78212-7200.

Notes

The authors declare no competing financial interest.

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